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The Crystal Structure of [2.2]Metaparacyclophane-1,9-diene

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[2.2]Metaparacyclophane-1,9-diene, $C_{16}H_{12}$, crystallizes in the orthorhombic system, space group $Pbca$, with $a = 14.91 \pm 1$, $b = 18.51 \pm 1$, $c = 8.133 \pm 5 \text{ \AA}$, $Z = 8$. The intensity data were measured with a four-circle diffractometer and scintillation counter. The structure was determined by symbolic addition and refined by block-diagonal least-squares analysis of 1378 reflexions to a final R index of 0.046. The two aromatic rings are inclined to each other at 41° . Both show significant boat distortion, which is moderate for the *meta*-bridged ring, and severe for the *para*-bridged ring.

[2.2]Metaparacyclophane-1,9-diene (I; Hylton & Boekelheide, 1968) is one of a series of compounds prepared by Professor Boekelheide and his associates. Nuclear magnetic resonance studies of the material in solution fail to resolve the individual protons of the *para*-bridged ring, suggesting that the molecule has $mm2$ symmetry, with the two rings perpendicular to each other (Boekelheide, 1968). This conformation could, of course, be simulated by rapid 'flipping' of the molecule between two conformations related by the apparent symmetry. The X-ray analysis was undertaken in order to determine the conformation in the solid state. It is found that the rings are in fact by no means perpendicular, but are inclined to each other at 41° .

Experimental

Crystal data

$C_{16}H_{12}$ F.W. 204.3

Orthorhombic,

$$a = 14.91 \pm 1, b = 18.51 \pm 1, c = 8.133 \pm 5 \text{ \AA}$$

(Cu $K\alpha_1$, $\lambda = 1.5405 \text{ \AA}$), $V = 2245 \text{ \AA}^3$, $D_m = 1.20 \text{ g.cm}^{-3}$, $Z = 8$, $D_x = 1.21 \text{ g.cm}^{-3}$, $\mu = 6.3 \text{ cm}^{-1}$.

Space group $Pbca$ (D_{2h}^{15}) (from precession photographs. Systematic absences: $0kl$ for k odd, $h0l$ for l odd, $hk0$ for h odd).

The crystals supplied were colourless and translucent, displaying various faces of which [010] was the most prominent. The specimen used for data collec-

tion was a triangular plate, 0.3 mm thick, and 0.5 mm to the side. The material was found to decompose in conditions of moderate humidity, and the specimen was therefore enclosed in a thin-walled Lindemann glass capillary.

The intensities were measured with a Picker four-circle diffractometer and scintillation counter, using nickel-filtered Cu $K\alpha$ radiation with pulse-height discrimination. The θ - 2θ scan method was used (2° for $2\theta < 100^\circ$, 3° otherwise), and background counts were measured at the beginning and end of each scan. Reflexions for which the net count was less than 10, or less than 20% of the gross count, were treated as unobserved. The yield of observed reflexions was rather low. In the range explored ($2\theta < 130^\circ$), 1380 of a possible 1904 reflexions were observed, and of these the net count exceeded four times the threshold value for only 956. Absorption corrections were not applied.

The structure was determined routinely by the symbolic addition procedure of Karle & Karle (1966). Refinement was by block-diagonal least-squares. The weighting scheme used was $w = w_1 w_2$, where

$$\begin{aligned} w_1 &= F_o/10 \text{ for } F_o < 10 \\ &= 10/F_o \text{ for } F_o \geq 10 \\ w_2 &= 2.5 \sin^2 \theta \text{ for } \sin^2 \theta < 0.4 \\ &= 1 \text{ for } \sin^2 \theta \geq 0.4 \end{aligned}$$

(The nominal minimum value of F_o is 2.5).

This weighting scheme was reasonably effective in removing trends in the magnitudes of the residuals.

The scattering factor curves used were those of Hanson, Herman, Lea & Skillman (1964). The hydrogen atoms were initially assumed to lie in chemically reasonable positions, and their parameters were refined concurrently with those of the carbon atoms. During refinement it was apparent that the stronger reflexions were suffering from extinction, and they were corrected in the manner described by Pinnock, Taylor & Lipson (1956). 65 reflexions were corrected, but the correction exceeded 10% of F_o for only 12 of these. The procedure was considered to be of questionable validity for the strongest reflexion (040; $F_c=277$), which was given zero weight in subsequent refinement. (It should be noted that the weighting scheme discriminates so strongly against the very intense reflexions that the form of the extinction correction is of little consequence to the analysis). Because of unacceptably poor

agreement one other reflexion was given zero weight: for 531, $F_o=5.7$, and $F_c=0$; the counts recorded for this reflexion presumably result from some extraneous phenomenon.

In the final cycle no coordinate shift exceeded 40% of the corresponding e.s.d. The final parameters are given in Table 1. Observed and calculated structure amplitudes are compared in Table 2. The agreement is satisfactory, and there can be no doubt of the essential correctness of the structure. The agreement residual is 0.046, for observed reflexions only. A three-dimensional difference synthesis is reasonably featureless; the residual electron density is everywhere within the limits $\pm 0.12 \text{ e.}\text{\AA}^{-3}$. (Peak electron densities range from 4.7 to 7.2 $\text{e.}\text{\AA}^{-3}$ for carbon atoms, and from 0.55 to 0.94 $\text{e.}\text{\AA}^{-3}$ for hydrogen atoms.)

Table 1. Final atomic parameters (e.s.d.'s)

The coordinates are fractional, $\times 10^4$. U_{ij} values are in $\text{\AA}^2 \times 10^4$. B values are in \AA^2 .

(a) Positions of the carbon atoms

The B values given are equivalent isotropic values in the directions of the principal axes of the vibration ellipsoids.

Atom	x	y	z	B_1	B_2	B_3
C(1)	5057(2)	4221(2)	3457(3)	10.1	9.1	4.4
C(2)	5754(2)	4105(2)	2467(4)	9.7	8.4	4.4
C(3)	5711(1)	3840(1)	793(3)	7.0	5.9	3.0
C(4)	6346(1)	3348(1)	130(4)	11.9	5.8	3.1
C(5)	6210(2)	3038(1)	-1376(4)	11.4	6.8	4.2
C(6)	5454(2)	3168(1)	-2263(3)	10.7	5.7	4.1
C(7)	4812(1)	3649(1)	-1686(3)	7.0	4.8	3.7
C(8)	4997(1)	4015(1)	-231(2)	5.3	4.6	2.9
C(9)	3920(2)	3274(1)	-2484(3)	9.4	6.3	4.3
C(10)	3142(2)	3831(1)	-1694(4)	10.1	6.8	4.2
C(11)	3108(1)	3877(1)	98(3)	9.0	5.8	3.0
C(12)	3075(1)	4541(1)	921(4)	10.3	5.4	3.6
C(13)	3529(2)	4646(1)	2307(4)	10.3	6.1	4.0
C(14)	4128(2)	4086(1)	2877(3)	8.7	6.8	3.6
C(15)	3901(2)	3385(1)	2383(3)	8.3	5.9	3.6
C(16)	3401(1)	3283(1)	1000(3)	8.6	5.6	3.0

(b) Anisotropic temperature-factor coefficients

The temperature factor is defined as

$$\exp\{-2\pi^2(U_{11}a^*2h^2 + 2U_{12}a^*b^*hk + 2U_{13}a^*c^*hl + U_{22}b^*2k^2 + 2U_{23}b^*c^*kl + U_{33}c^*2l^2)\}.$$

Atom	U_{11}	U_{12}	U_{13}	U_{22}	U_{23}	U_{33}
C(1)	1213(20)	-90(17)	-177(14)	1092(17)	-178(14)	685(14)
C(2)	794(14)	-12(13)	-268(13)	1050(17)	-28(16)	994(16)
C(3)	452(9)	-146(8)	10(10)	708(10)	73(10)	858(13)
C(4)	419(9)	-45(9)	93(12)	882(14)	316(16)	1340(21)
C(5)	751(14)	7(12)	390(14)	856(16)	-2(14)	1224(20)
C(6)	994(16)	-120(12)	390(13)	762(12)	-105(12)	842(15)
C(7)	734(11)	-113(9)	166(10)	613(10)	-6(10)	627(11)
C(8)	479(8)	-97(7)	95(8)	524(9)	18(8)	624(10)
C(9)	1115(17)	-81(13)	-187(13)	786(12)	-38(12)	627(11)
C(10)	752(14)	-38(11)	-328(14)	842(16)	-82(13)	1078(17)
C(11)	398(8)	-84(8)	24(10)	715(10)	-28(12)	1142(17)
C(12)	511(9)	46(9)	200(12)	674(10)	-18(13)	1257(19)
C(13)	811(14)	4(11)	342(14)	759(12)	-175(13)	1012(18)
C(14)	940(15)	-55(12)	248(11)	892(14)	-71(11)	594(12)
C(15)	796(12)	-42(10)	272(11)	708(12)	94(11)	744(12)
C(16)	533(9)	-133(8)	205(10)	632(10)	25(11)	1008(16)

(c) Hydrogen atom positions and isotropic temperature factors

Atom	x	y	z	B
H(1)	5146(23)	4364(19)	4514(44)	7.6(0.8)
H(2)	6346(23)	4186(17)	2919(40)	7.3(0.8)
H(4)	6850(20)	3233(16)	810(37)	6.4(0.7)
H(5)	6689(22)	2660(19)	-1856(44)	8.1(0.9)
H(6)	5354(18)	2926(15)	-3363(36)	5.9(0.6)
H(8)	4585(11)	4402(9)	90(22)	1.5(0.3)
H(9)	3912(21)	3671(17)	-3680(40)	7.0(0.7)
H(10)	2579(21)	3916(15)	-2352(36)	6.3(0.6)
H(12)	2810(15)	4946(13)	358(31)	4.5(0.5)
H(13)	3649(19)	5184(14)	2683(34)	5.6(0.6)
H(15)	4219(14)	2973(12)	2882(27)	3.7(0.5)
H(16)	3385(15)	2767(12)	508(30)	4.0(0.5)

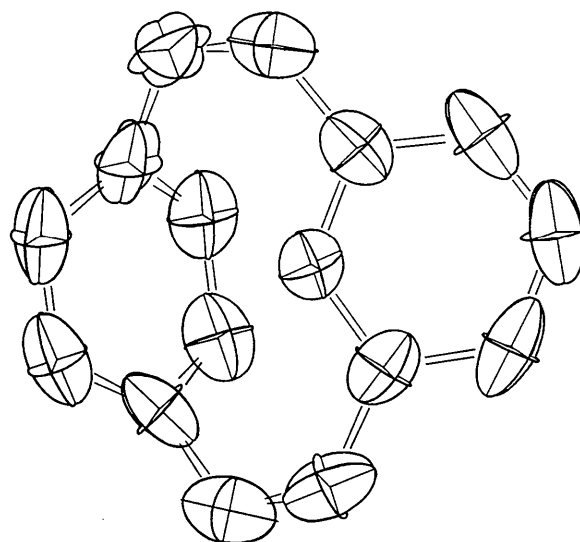
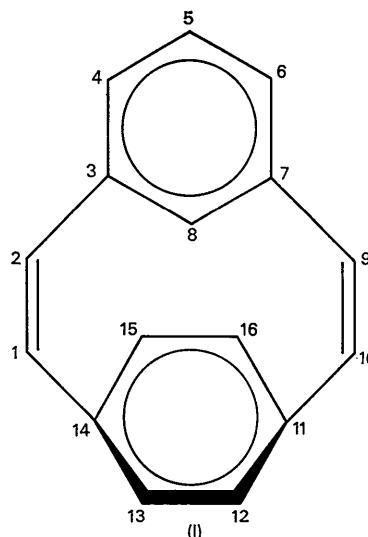


Fig. 1. The principal planes of the thermal motion ellipsoids of 50% probability.

Table 2. Observed and calculated structure amplitudes, x 10

An asterisk denotes the threshold value of an unobserved reflexion.

Table with multiple columns containing numerical data for observed and calculated structure amplitudes. The table is organized into several vertical sections, each with its own set of column headers. Asterisks are used to denote threshold values for unobserved reflexions.

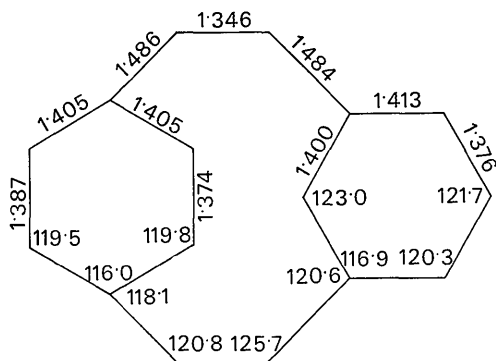
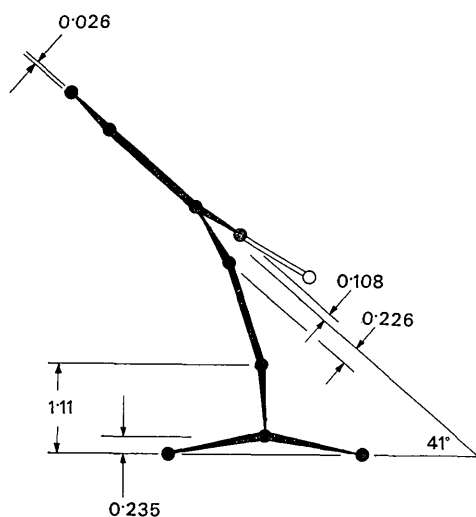
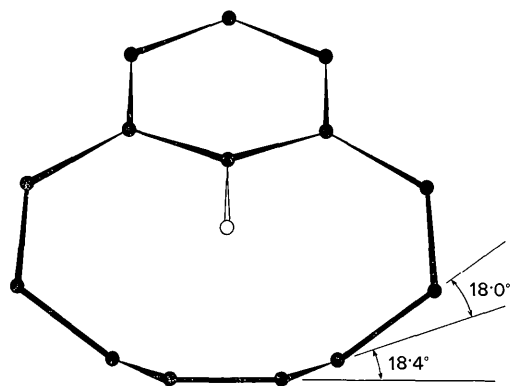


Fig. 2. Mean bond lengths (Å) and angles (°) assuming mirror symmetry.



(a)



(b)

Fig. 3. Two views of the slightly idealized carbon skeleton. Distances shown are in Å. The open circle indicates the position of H(8). The directions of view are: (a) the normal to the approximate mirror plane; (b) the intersection of the mirror plane with the mean plane of the *para*-bridged ring.

Table 3 (cont.)

Mean aromatic bond distance	1.396 Å
Mean double-bond distance	1.346
Mean single-bond distance	1.485
Mean C-H bond distance	0.99
	(0.91 to 1.07)

Table 4. Comparison of bond angles

Angle	Value	e.s.d.	Mean
C(14)—C(1)—C(2)	120.5°	0.3°	120.8°
C(9)—C(10)—C(11)	121.1	0.2	
C(1)—C(2)—C(3)	126.1	0.3	125.8
C(7)—C(9)—C(10)	125.4	0.2	
C(2)—C(3)—C(4)	122.6	0.2	122.3
C(6)—C(7)—C(9)	122.0	0.2	
C(2)—C(3)—C(8)	121.0	0.2	120.6
C(8)—C(7)—C(9)	120.2	0.2	
C(4)—C(3)—C(8)	116.3	0.2	116.9
C(6)—C(7)—C(8)	117.5	0.2	
C(3)—C(4)—C(5)	120.2	0.2	120.3
C(5)—C(6)—C(7)	120.4	0.2	
C(4)—C(5)—C(6)	121.7	0.2	
C(3)—C(8)—C(7)	123.0	0.2	
C(10)—C(11)—C(12)	121.9	0.2	121.6
C(13)—C(14)—C(1)	121.3	0.2	
C(10)—C(11)—C(16)	118.2	0.2	118.1
C(15)—C(14)—C(1)	118.0	0.2	
C(12)—C(11)—C(16)	116.9	0.2	117.0
C(13)—C(14)—C(15)	117.1	0.2	
C(11)—C(12)—C(13)	119.7	0.2	119.5
C(12)—C(13)—C(14)	119.3	0.2	
C(14)—C(15)—C(16)	119.8	0.2	119.8
C(15)—C(16)—C(11)	119.9	0.2	

Mean C—C—H angle

119
(116 to 123)

Discussion

The mean bond lengths and angles given in Table 3 and Fig. 2 are consistent with expectation and require no comment. Views of the molecule are given in Fig. 3. It is clear that whatever may be the case in solution, the molecule in the crystal does not have the predicted *mm*2 symmetry. Far from being mutually perpendicular, the aromatic rings are inclined to each other at 41°.

Significant strain may be inferred from the distortion from planarity of the aromatic rings and their substituents. The distortion of the *meta*-bridged ring is comparatively mild: the displacements noted in Fig. 3(a) (0.026, 0.108 and 0.226 Å) may be compared with the corresponding displacements of 0.042, 0.143 and 0.41 Å for [2.2]metacyclophane (Brown, 1953). The distortion of the *para*-bridged ring is much more severe: the displacements of the bridgehead atoms and their substituents from the mean plane of the unsubstituted atoms are 0.235 and 1.11 Å. The distortion may also be expressed in terms of the angles α and β , as in Fig. 3(b). The values given there (18.4 and 18.0°) may be compared with the corresponding values of 14 and 15° found for the highly strained compound [2.2]paracyclophane-1,9-diene (Coulter & Trueblood, 1963).

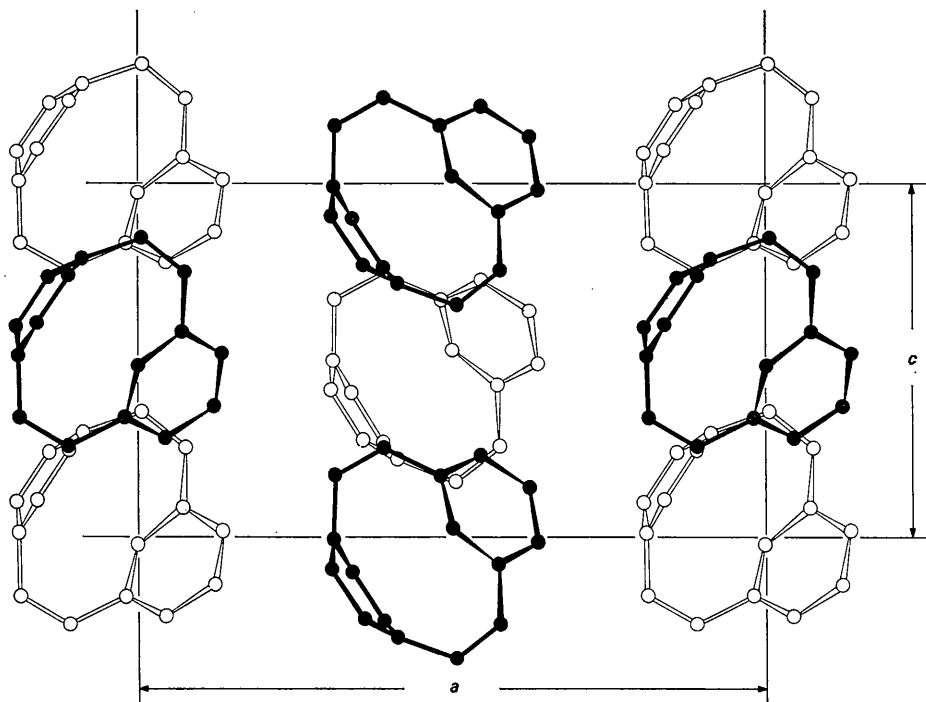


Fig. 4. A projection along b of that part of the structure lying between $y=0$ and $y=\frac{1}{2}$. The remainder is related to the part shown by a centre of symmetry at $\frac{1}{2}, \frac{1}{2}, \frac{1}{2}$.

In spite of this distortion C(8) lies only 2.71 Å, and H(8) only 2.16 Å, from the plane of the unsubstituted atoms of the *para*-bridged ring. If the predicted $mm2$ symmetry were to be adopted without further distortion, simply by constraining the *meta*-bridged ring and the bridging atoms to lie in a plane perpendicular to that of the *para*-bridged ring, these distances would be reduced to 2.57 and 1.59 Å. It seems certain that such a close approach would have to be avoided by further severe distortion of the molecule. Whether the energy required to cause such distortion would be sufficient to prevent significant 'flipping' of the molecule (in solution) is an open question.

[A distance of 0.98 Å has been assumed for C(8)–H(8). This is the distance actually observed; moreover it is in good agreement with C–H distances observed in many X-ray structure analyses. It does not represent internuclear distance, but rather the distance between centres of gravity of electron-density distributions. It is of course the electron-density distribution which is relevant in consideration of non-bonded contact].

The packing of the molecules is illustrated in Fig. 4. The shortest intermolecular contacts are: C–C, 3.59; C–H, 2.79; H–H, 2.46 Å.

The problem was suggested, and the specimen material supplied by Professor V. Boekelheide. Dr C. K. Johnson's (1965) *ORTEP* program was used for Fig. 1. All other computations (including the phasing of the reflexions) were carried out using the programs of Ahmed, Hall, Pippy & Huber (1966). The assistance of those named is gratefully acknowledged.

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